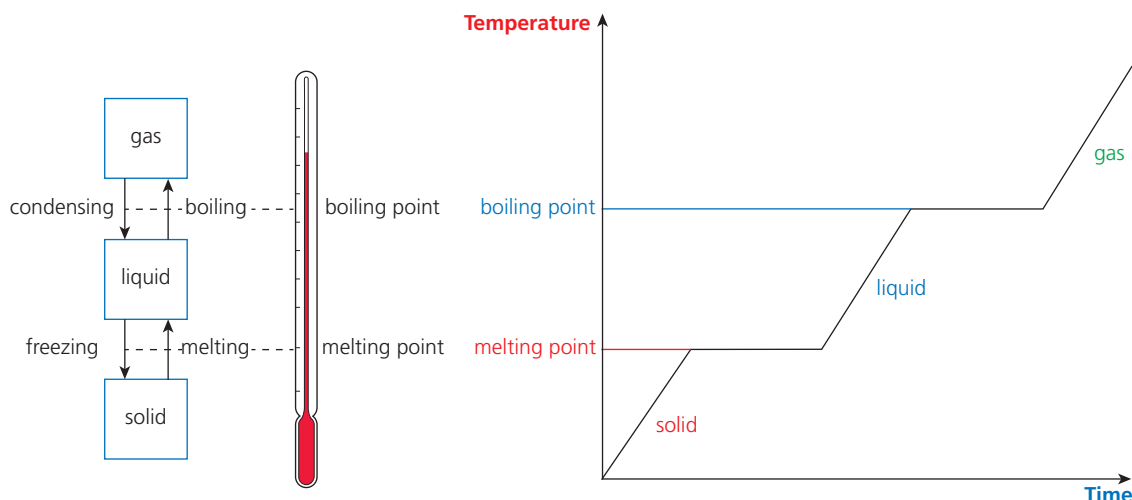


▼ Melting point and boiling point



▼ Boiling and evaporation

	Boiling	Evaporation
State change	Change from liquid to gas	Change from liquid to gas
Temperature	Only at the boiling point	Any temperature below the boiling point
Location	Throughout the liquid Bubbles of gas form throughout the substance	Only from the surface of the liquid Particles gain enough energy to break away from the surface

C1.4 Gas pressure

	Question	Answer
1	What is kinetic theory?	The idea that all particles are moving
2	What causes gas pressure?	Gas particles colliding with the surface of a container
3	What three factors affect gas pressure?	Temperature, size of the container, number of gas particles
4	What is the relationship between temperature and gas pressure?	As temperature increases, the gas pressure increases
5	What is the relationship between the size of a container and gas pressure?	As the size of the container decreases, the gas pressure increases
6	What is the relationship between gas pressure and the number of gas particles?	As the number of gas particles increases, the gas pressure increases
7	What happens to particles when you heat them up?	They move faster